

# ATOMS REVIEW QUESTIONS

1. The number of protons & neutrons for a specific isotope is called
2. Calculate the neutrons an atom with an atomic mass of 11, atomic number of 5 & 5 electrons.
3. In a neutral atom, the electrons must equal
4. The number of protons identifies
5. List a neutrons charge, mass, & location
6. List a protons charge, mass, location
7. To find the # of protons of an atom, you
8. List an electron's charge, mass, location
9. How many electrons can be held in the 1st energy level?
10. Name the scientist who proved the existence of an atom.
11. Name the Greek philosopher that predict atoms exist.
12. Name the scientist who discovered electrons in orbits around the nucleus.
13. Isotopes are elements with
14. Name the scientist who discovered electrons.
15. This scientist discovered the nucleus was a tiny, dense, positive area of an atom in which the rest is empty space?
16. An atom's mass is mainly found in the
17. In the electron cloud, electrons are arranged in
18. Some of the info in chemistry is gathered by not directly seeing or touching. This is called
19. This identifies an element.
20. The location of an electron is most likely determined by
21. The coefficients for this equation are  $\_Na + \_O_2 \rightarrow \_Na_2O$
22. What is the difference b/w atomic mass & mass number?
23. The coefficients for this equation are  $\_C + \_H_2 \rightarrow \_CH_4$
24. The coefficients for this equation are  $\_H_2 + \_O_2 \rightarrow \_H_2O$
25. How many electrons can the first 3 energy levels hold?
26. Identify the energy levels for Na (atomic # 11).
27. Identify the energy levels for O (atomic # 8).
28. How many energy levels are occupied in an atom of N (atomic # 7)?
29. How many energy levels are occupied in an atom of Al (atomic # 13)?
30. Calculate the # of neutrons in an atom of F (at # 9; mass # 19).

## ATOMS REVIEW answers

1. mass #
2. 6
3. protons
4. atomic #
5. 0, 1, nucleus
6. +, 1, nucleus
7. subtract neutrons from atomic mass
8. -, .0006, electron cloud
9. 2
10. John Dalton
11. Democritus
12. Niels Bohr
13. same # protons; different # neutrons
14. JJ Thomson
15. Ernest Rutherford
16. nucleus
17. energy levels
18. indirect evidence
19. atomic number
20. how much energy it has
21. 4,0,2
22. atomic mass is an average of all isotopes; mass # is mass of 1 isotope
23. 0,2,0
24. 2,0,2
25. 2-8-18
26. 2-8-1
27. 2-6-0
28. 2
29. 3
30. 10